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# **Electrochemical molecularly imprinted polymers as material for pollutant detection**

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Keywords: electrochemistry, molecularly imprinted polymer, ferrocene, benzo(a)pyrene

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**ABSTRACT:** New electrochemical molecularly imprinted polymers (e-MIPs) for benzo(a)pyrene (BaP) based on vinylferrocene (VFc) as redox tracer directly integrated into binding cavities during polymerization with crosslinker are reported. Impacts on the BaP detection are discussed according to: 1) the addition or no of 4-vinylpyridine (4VP) as comonomer, 2) the use of ethylene glycol dimethacrylate (EDMA) or divinylbenzene (DVB) as crosslinker.

After characterization, these new e-MIPs and their corresponding non-imprinted polymers (e-NIPs) were incorporated in carbon paste electrodes and measurements were carried out by square wave voltammetry. All analyses revealed that e-MIP based on VFc and EDMA is the best to detect the template, with an imprinting factor of 7.3 and a LOD of 0.93  $\mu\text{mol/L}$  of BaP.

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## 1. Introduction

Research on molecularly imprinted polymers (MIPs) is currently a topic of intensive investigations due to their recognition properties, their easy synthesis and their high stability [1,2]. Such polymers are frequently described as "plastic antibodies" due to their high selectivity [3]. MIPs are prepared by copolymerization of a crosslinker and a functional monomer interacting with a template [4,5]. The removal of the template from the polymer network generates complementary sites that have specific interactions with the template and its analogs.

In this context, MIPs are of great interest for development of sensors due to their high potential as stable recognition elements, and in particular in electrochemical sensors [6,7]. Since few years, we develop electrochemical molecularly imprinted polymers (e-MIPs) with a redox tracer as functional monomer directly integrated inside binding cavities during the copolymerization with the crosslinker. By this way, redox binding cavities are obtained and can be used for sensing applications [8]. Initially, this new strategy was developed to detect benzo(a)pyrene (BaP) by the introduction of a ferrocenyl monomer whose cyclopentadienyl rings could generate aromatic stacking interactions with BaP [9]. Recently, we extended this strategy to e-MIP for bisphenol A (BPA) and integrated these new e-MIPs in screen printed electrodes. A limit of detection of 0.06 nmol/L of BPA was obtained [10,11]. This strategy has also been taken up to detect vancomycin, an antibiotic target, with electroactive MIPs prepared by solid-phase synthesis with different amounts of VFc and ferrocenylmethyl

methacrylate into polymerization mixture [12]. Recently the use of supercritical carbon dioxide as porogenic solvent has been reported to prepare ferrocenyl-based MIP for BPA [13]. By this way authors indicate a quantification of BPA in the range of 4.7-8 nmol/L.

To increase the performances of the electrochemical e-MIPs for BaP detection, we describe here the evolution of the system based on the use of VFc and 4-vinylpyridine (4VP) as comonomer and ethylene glycol dimethacrylate (EDMA) or divinylbenzene (DVB) as crosslinker. 4VP monomer and DVB crosslinker were reported by Lai et al. [14] to increase the performance of MIPs in terms of imprinting factor, selectivity and adsorption capacity via  $\pi$ - $\pi$  interactions between the aromatic rings present in these compounds and in BaP. Moreover, DVB is known to give homogeneous spherical particles by precipitation polymerization which can present an advantage to prepare carbon paste electrode [15]. To evaluate the influence of 4VP as functional comonomer and of the aromatic crosslinker DVB on e-MIP properties, three new e-MIPs (and corresponding non imprinted polymers e-NIPs) were prepared using a hybridation oven: 1) VFc as monomer with DVB as crosslinker (VFc-DVB), 2) VFc and 4VP as comonomer with EDMA as crosslinker (VFc-4VP-EDMA) and 3) VFc and 4VP as comonomer and DVB as crosslinker (VFc-4VP-DVB). We present in this article the synthesis, characterization and electrochemical properties of these new copolymers compared to the e-MIP VFc-EDMA prepared in the same conditions.

## **2. Method and materials.**

### **2.1. Reagents and chemicals**

Vinylferrocene (VFc), 4-vinylpyridine (4-VP), ethylene glycol dimethacrylate (EDMA), divinylbenzene (DVB), acetylferrocene, anthracene, fluorene, phenanthrene and tetrabutylammonium hexafluorophosphate were purchased from Acros. Benzo(a)pyrene (BaP), chrysene, fluoranthene, pyrene, dibenzo(a,h)anthracene, tetrabutylammonium hexafluorophosphate (TBAFP) were purchased from Aldrich and Azobisisobutyronitrile (AIBN) from Fluka. Toluene, heptane and acetonitrile were analytical grade and purchased from VWR and used without any further purification. Ultrapure water (Millipore) was used for analysis. The quality of acetonitrile was HPLC grade.

### **2.2. Synthesis**

Polymerization by precipitation was carried out in an UVP HB-100 hybridation oven with molar ratios summarized in Table 1. Functional monomer, crosslinker agent and BaP (in the case of MIPs) were introduced in the Pyrex glass tube (diameter of 35 mm and volume of 120

mL) under nitrogen gas and kept in the refrigerator at 4°C for 30 minutes to promote the formation of functional monomer complex-BaP. Then, the solvent was added (monomer/solvent ratio is 3 % by weight) and the whole mixture was degassed for 15 minutes in an ice-water bath. After the treatment the radical initiator (AIBN) was added (the ratio AIBN monomer is 2 % by weight) and again the mixture was degassed for another 5 minutes under similar conditions. The reaction medium was heated in at 8 rpm along the axis of the tube. The temperature of the oven was increased by 5 °C every 15 minutes from 30 °C to 80 °C and kept constant at this temperature until the end of the polymerization, at least 20 hours. The polymer was recovered by centrifugation at 10,000 rpm and washed with a mixture of acetonitrile and toluene except for the VFc-4VP-EDMA polymers for which a mixture of heptane and toluene was used. In fact for Fc-4VP-EDMA it was impossible to have precipitates with a suitable yield (more 10 %) using the acetonitrile-toluene mixture. Elemental analyses, and particularly iron dosage, were made by the Central Analysis Service of the CNRS Laboratory of Solaize USR- 59C/CNRS and confirm the well integrations of monomers (Table 1).

Table 1. Polymerization by precipitation results with the molar ratios for e-MIPs: BaP/VFc/EDMA: 1/5/25; BaP/VFc/4VP/EDMA: 1/2.5/2.5/25; BaP/VFc/DVB: 1/5/25 and BaP/VFc/4VP/DVB: 1/2.5/2.5/25. Identical conditions were used for e-NIPs without BaP.

Copolymer	BaP (mmol)	Quantities (mmol)			Solvent %v/v		Solvent (mL)	Yield (%)	Dn <sup>a</sup> (nm)	CV <sup>b</sup> (%)	S <sub>BET</sub> (m <sup>2</sup> /g)	Vp (cm <sup>3</sup> /g)	Elemental analysis		
		VFc	4-VP	EDMA or DVB	Acetonitrile (ACN) or Heptane (HP)	Toluene							αVFc (% mol)	α4VP (% mol)	αEDMA or αDVB (% mol)
e-MIP VFc-EDMA	0.32	1.60	0	8	70 (ACN)	30	100	59	580 ± 170	29	14.6	0.022	19	/	81 (83) <sup>d</sup>
e-NIP VFc-EDMA	0	1.60	0	8	70 (ACN)	30	100	80	210 ± 60	29	43.8	0.153	18	/	82 (83) <sup>d</sup>
e-MIP VFc-4VP-EDMA	0.24	0.60	0.6	6	95 (HP)	5	60	40	nd	nd	241.5	0.132	8	8	84 (84) <sup>d</sup>
e-NIP VFc-4VP-EDMA	0	0.60	0.6	6	95 (HP)	5	60	40	nd	nd	93.7	0.225	8	8	84 (84) <sup>d</sup>
e-MIP VFc-DVB	0.35	1.75	0	8.75	95 (ACN)	5	70	30	220 ± 40	19	125.1	0.258	9	/	91 (83) <sup>d</sup>
e-NIP VFc-DVB	0	1.75	0	8.75	95 (ACN)	5	70	40	250 ± 90	37	nd	nd	9	/	91 (83) <sup>d</sup>
e-MIP VFc-4VP-DVB	0.40	1	1	10	95 (ACN)	5	70	26	300 ± 150	49	298.2	0.161	2	11	87 (84) <sup>d</sup>
e-NIP VFc-4VP-DVB	0	1	1	10	95 (ACN)	5	70	28	570 ± 240	43	nd	nd	3	11	86 (84) <sup>d</sup>

<sup>a</sup> The average particle diameter (Dn) is determined from SEM images taken at 100 particles and using SmarTiff software.

<sup>b</sup> Coefficient of variation (CV) is defined by the average particle size of the standard deviation (CV% = Dn x 100/σ)

nd = not determined

<sup>d</sup> Theoretical values

E-MIPs and e-NIPs were analyzed by scanning electron microscopy (SEM) with Gemini® Supra 40VP apparatus. Small amount of polymer was deposited on a SEM support and metalized with gold. The images were registered under the following conditions: 3 kV, secondary electron detector, diaphragm 30. Average particle diameters ( $D_n$ ) were determined from SEM images taken at 100 particles and using SmarTiff software and coefficient of variation (CV) defined by the average particle size of the standard deviation (equation 1).  $D_n$  and CV values are summarized in Table 1.

$$CV = \frac{\text{Standard deviation}}{\text{average diameter}} \times 100 = \frac{\left( \sqrt{\frac{\sum (d - D_n)^2}{N}} \right)}{D_n} \times 100 \quad [\text{Equation 1}]$$

Brunauer–Emmett–Teller (BET) analyses were carried on a Gemini 2380 apparatus from Micrometrics.

Thermal analyzes were performed on a SDT Q600 from TA Instruments and thermograms were treated from the Universal Analysis software (TA Instruments). The heat made program consists of two steps: 1) an isotherm of 10 minutes at 25 °C to stabilize the measuring cell and 2) a ramp from 25 °C to 900 °C at a rate of 20 °C/min under nitrogen.

### 2.3. Quantification of BaP by HPLC-UV

All separations were achieved on using an isocratic system when only BaP was present or a gradient system for the mixtures of several Polycyclic Aromatic Hydrocarbons (PAHs) (fluorene, fluoranthene, anthracene, phenanthrene, chrysene, pyrene, and dibenzo(a,h)anthracene). The former consisted of pure acetonitrile while the latter consisted of a mixture of acetonitrile-water starting with 55 % acetonitrile for 10 minutes, then ascending linearly to 100% acetonitrile within 25 minutes, and being maintained for 20 minutes before finally dropped down to 55 % within 5 minutes. In each case, the flow rate was fixed at 0.5 mL/min and a 50 µL sample volume was injected into the chromatographic system.

Calibration curves at  $\lambda = 296$  nm were obtained using a series of standard solutions containing mixture of PAHs at five different concentrations. Three replicated injections for each concentration were made to ensure accurate and reproducible responses. Linearity was evaluated by using least square method. The theoretical limits of detection (LOD) at a signal/noise ratio of 3 and theoretical limits of quantification (LOQ) at a signal/noise ratio of

10 for all PAHs were in the range of 0.04-0.79 and 0.13-2.64 µg/L, respectively.

## 2.4. Batch binding experiments

10 mg of polymer were placed in a 30 mL glass vial and mixed with 10 mL of BaP solution of different concentrations varying from 0.01 to 10 mg/L. BaP solutions were prepared in water/acetonitrile (99/1, v/v). Then the vial was shaken under horizontal stirring at 200 rpm at room temperature until equilibrium (24 hours). The supernatants were collected. The amounts of adsorbed BaP were calculated by subtracting the supernatant concentrations from the initial concentrations of BaP in solution (Equation 2). The imprinting factor was determined as the ratio between the maximum binding capacity of the e-MIP and that of the e-NIP (Equation 3). The binding isotherm experiment was carried out in duplicate.

$$Q = \frac{\text{adsorbed quantity (mg or } \mu\text{mol)}}{\text{polymer weight (g)}} \quad [\text{Equation 2}]$$

$$\text{or } Q = \frac{(\text{Initial concentration} - \text{supernatant concentration}) \times \text{volume}}{\text{polymer weight (g)}}$$

$$\text{IF} = \frac{Q_{e-MIP}}{Q_{e-NIP}} \quad [\text{Equation 3}]$$

## 2.5. Cross-selectivity experiments

The binding isotherm experiments were carried out in duplicate. For cross-selectivity, experiment 10 mg of polymer were placed in a 30 mL glass vial and mixed with 10 mL of solution containing BaP and one interfering PAH. The studied PAHs used were fluorene, fluoranthene, anthracene, phenanthrene, chrysene, pyrene, and dibenzo(a,h)anthracene. The concentrations of BaP and the interfering PAH were 0.4 µg/L each. The vial was then shaken horizontally at 200 rpm at room temperature for 24 hours. The supernatant was analyzed by HPLC-UV in order to determine the equilibrium concentration of the target ( $C_e$ ) and the amount of analyte adsorbed by the polymer  $Q_e$  (Equation 4) in mg/g, with  $C_0$  = initial concentration,  $V$  = volume of solution and  $m$  = mass of polymer).

$$Q_e = \frac{(C_0 - C_e) \times V}{m} \quad [\text{Equation 4}]$$



The selectivity coefficient is calculated from Equation 5.

$$k = \frac{Q_e(BaP) * C_e(PAH)}{Q_e(PAH) * C_e(BaP)} \quad \text{[Equation 5]}$$

Finally, the selectivity of e-MIP can be compared to the e-NIP by the relative selectivity coefficient ( $k'$ ) which is the ratio of the selectivity coefficient ( $k$ ) of a MIP to that of the corresponding e-NIP:  $k' = \frac{k_{MIP}}{k_{NIP}}$

## 2.6. Electrochemistry analyses

For electrochemistry analyses e-MIP and e-NIP were mixed with a carbon paste (Gwent group, C10903P14) (1/4, w/w). Mixtures were homogenized with a small amount of dichloromethane and filled in a hollow working 5 mm diameter electrode (Metrohm, RDE.BLANK 80184). The electrode was then dried under vacuum at 30 °C overnight and the surface was polished with a sand paper. The dry polymer and carbon paste weighed 11.4 mg. The working electrode containing polymer was then placed in an electrolytic cell in presence of a reference electrode (Hg/Hg<sub>2</sub>Cl<sub>2</sub>/KCl sat) and a counter electrode (Pt, 1 mm diameter). The electrolyte used was 0.1M tetrabutylammonium hexafluorophosphate in acetonitrile. The solution was purged 5 minutes with an argon flow. The blank measurement was registered after 30 minutes of electrode submersion in order to get a stabilized signal. Then 100 µL of BaP solution at 0.2 mg/L was added into the electrolyte solution. Square Wave Voltammetry (SWV) parameters *vs* SCE are: Initial potential (EI): 0.0 V; Vertex potential (EV): 1.0 V; Final potential (EF): 0.0 V; Pulse (PH/PW): 0.05 V; Pulse duration: 50 ms; Pulse height (SH): 5 mV; Scan rate: 20 mV/s.

## 3. Results

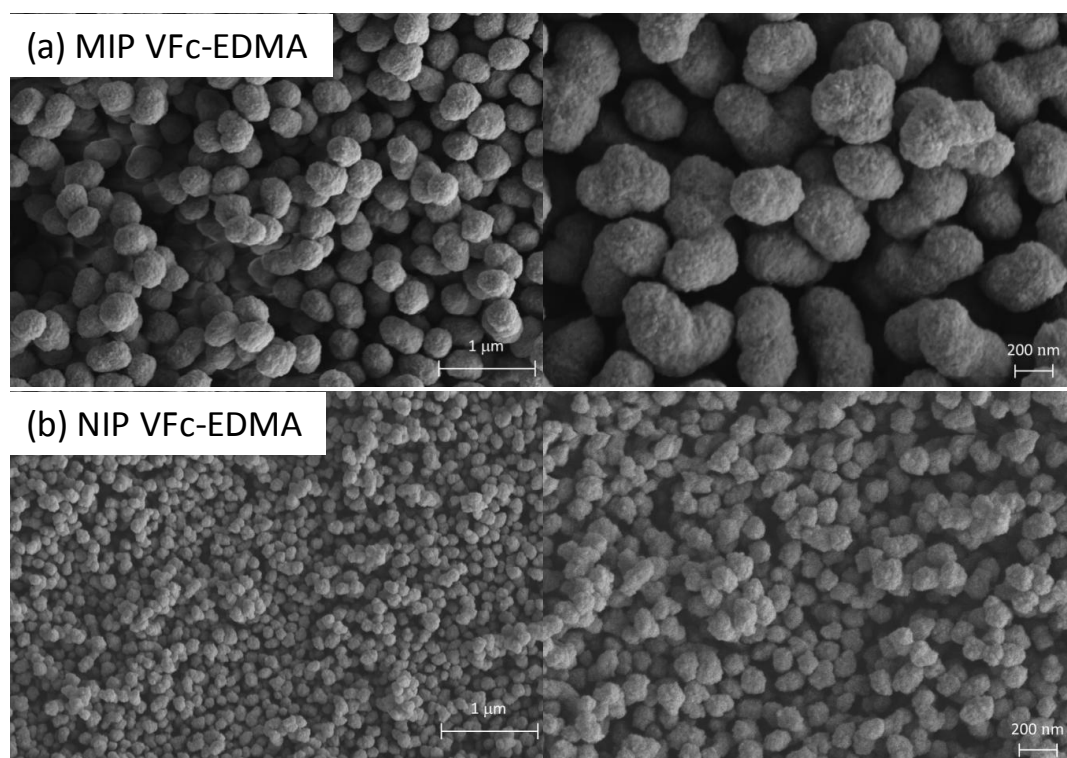
### 3.1 Synthesis and characterizations of e-MIPs and e-NIPs

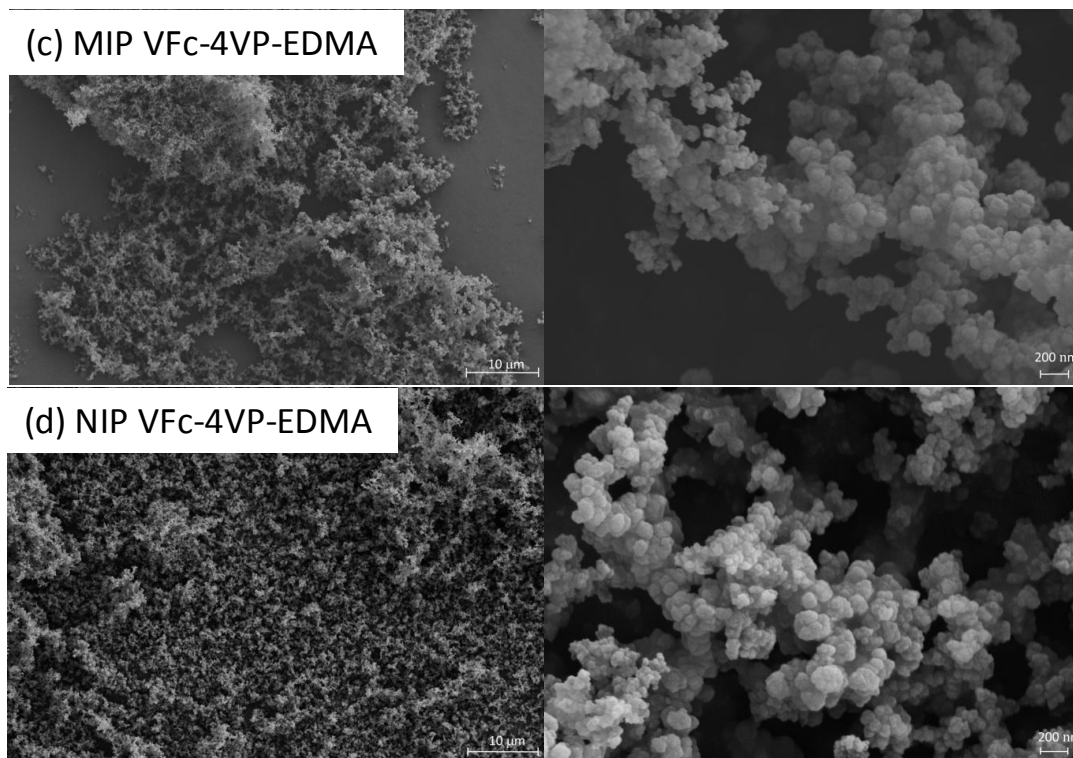
Particles of e-MIPs and e-NIPs were prepared by precipitation polymerization. This convenient process requires to use a solvent, or mixture of solvents, in which the monomers, the initiator and the template molecule are initially soluble. To form particles, the choice of the solvent is primordial since it must allow the precipitation of growing chains. In order to achieve this, mixtures of acetonitrile/toluene were used for VFc-EDMA and DVB based polymers (70/30 and 95/5 respectively). However, as far as VFc, 4VP and EDMA monomers

were concerned, the precipitation could not be resolved in any mixture based on these two solvents, regardless of their relative proportions. This was the reason why acetonitrile had to be replaced by heptane.

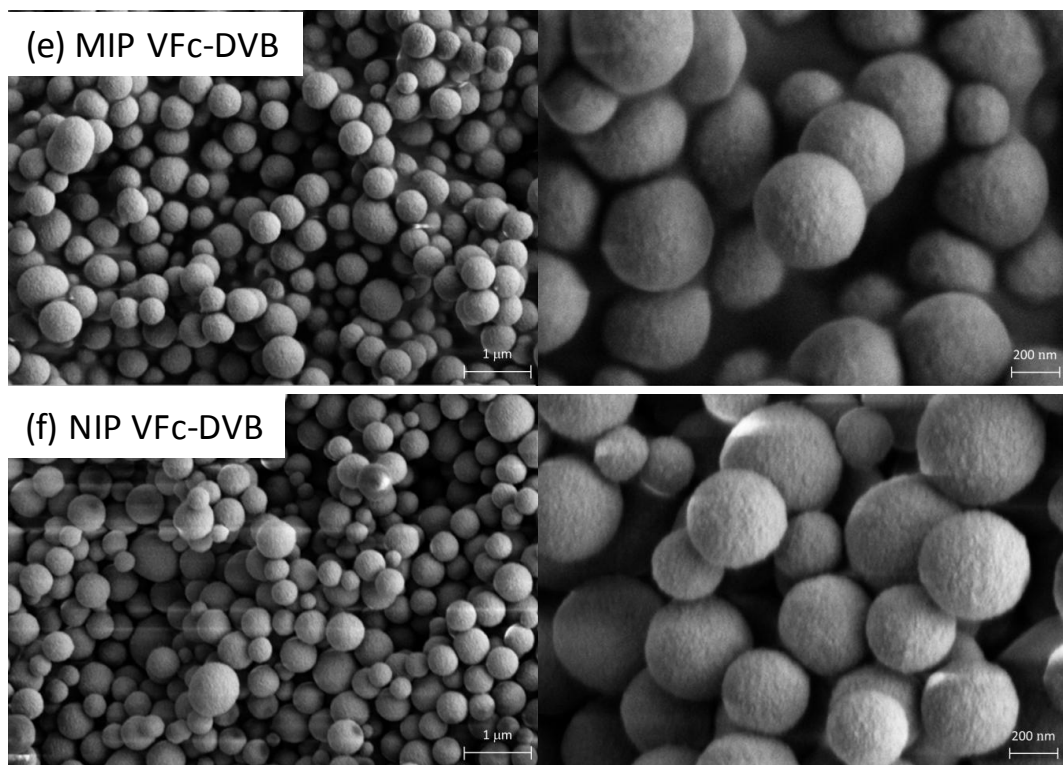
SEM pictures revealed that, for polymers with DVB as crosslinker, well defined beads with diameters were obtained with sizes varying from 220 to 570 nm (Figure 1). The use of EDMA as crosslinker for VFc-EDMA system led to aggregated beads in the similar size range (210 to 580 nm), whereas the particles were poorly defined for VFc-4VP-EDMA system. This appears to be a drastic effect of the change of monomers and solvent (introduction of heptane) for the obtaining of particles in a spherical form via precipitation polymerization.

Elemental analyses for copolymers based on EDMA were consistent with theoretical values which confirmed the good incorporation of VFc and 4VP monomers (Table 1). On the contrary with DVB, the results showed a difference between theoretical and experimental values revealing that the VFc was not properly incorporated. George and Hayes [16] have reported that VFc lead to a degree of polymerization independent on the concentration of the initiator or a poor incorporation of this monomer with styrene. These authors did the hypothesis of an intramolecular electron transfer from the iron atom to the radical center generating an anionic end which leads to an intramolecular termination reaction [17]. This hypothesis is consistent with our results.





**Figure 1.** SEM pictures of (a) e-MIP VFc-EDMA, (b) e-NIP VFc-EDMA, (c) e-MIP VFc-4VP-EDMA, (d) e-NIP VFc-4VP-EDMA.



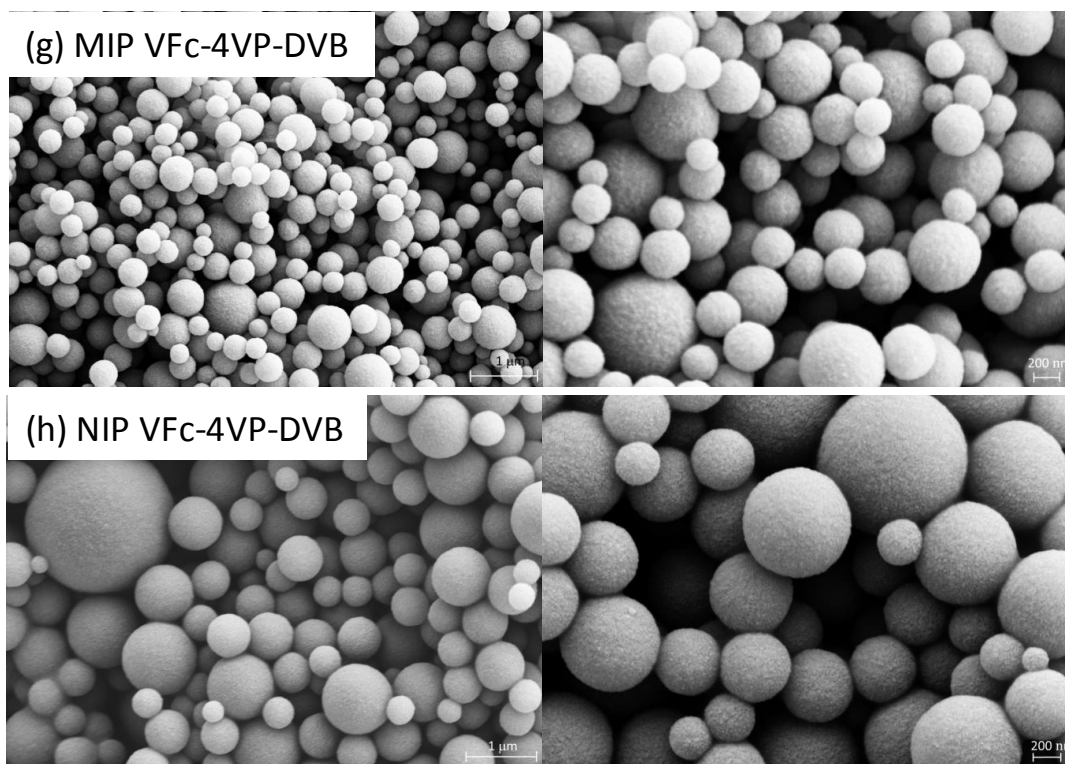


Figure 1. SEM pictures of (e) e-MIP VFc-DVB, (f) e-NIP VFc-DVB, (g) e-MIP VFc-4VP-DVB, (h) e-NIP VFc-4VP-DVB.

Thermogravimetric analyzes under nitrogen at 20 °C/min indicated a similar behavior for all copolymers. The thermal degradations of VFc-EDMA and VFc-4VP-EDMA copolymers began at 250 °C. DVB-based copolymers were more resistant with a beginning of degradation at 320 °C as expected for polymers integrating aromatic groups (Figure 2).

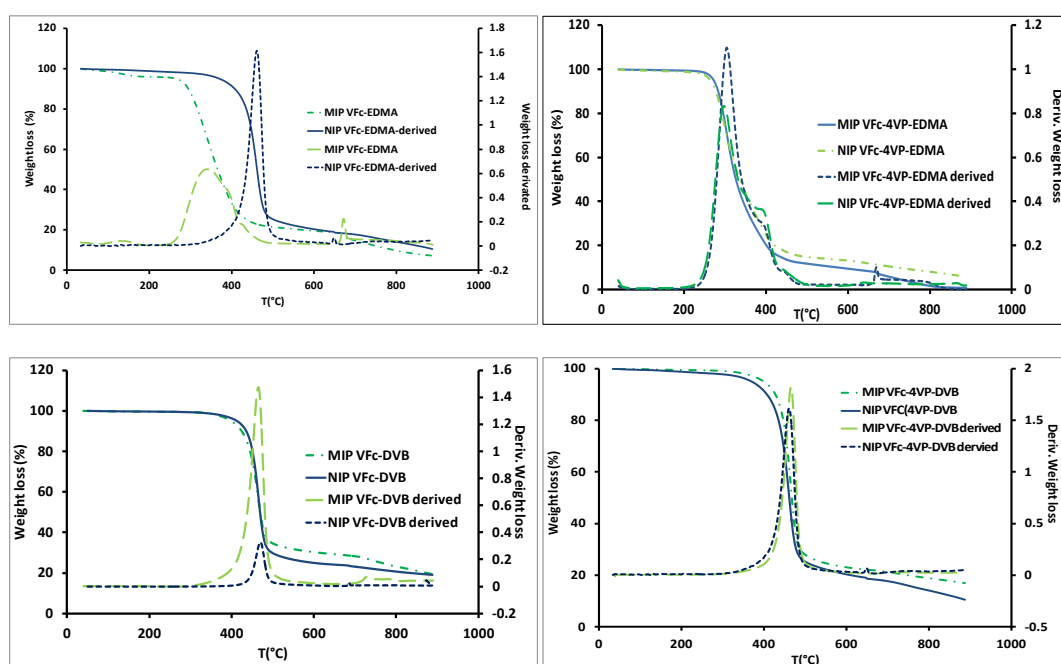


Figure 2. TGA curves in nitrogen with a heating rate of 20 °C/min for e-MIP/e-NIP for VFc-EDMA, VFc-4VP-EDMA, VFc-DVB and VFc-4VP-DVB.

BET analyses were carried out for all copolymers (Table 1, Figure 3). Isothermal adsorption/desorption of nitrogen for VFc-EDMA was a type I isotherm revealing that this copolymer is non porous confirmed by the close values of surface area determined experimentally. Other copolymers had type II isotherms which are characteristic of nonporous or macroporous materials. The comparison of surface area determinates from BET measurements seemed to indicate that addition of 4VP as functional comonomer increases the specific surface of copolymers. However the interpretation of these data was not easy because porosity variations may be linked to porogen solvent or monomer incorporation rate.

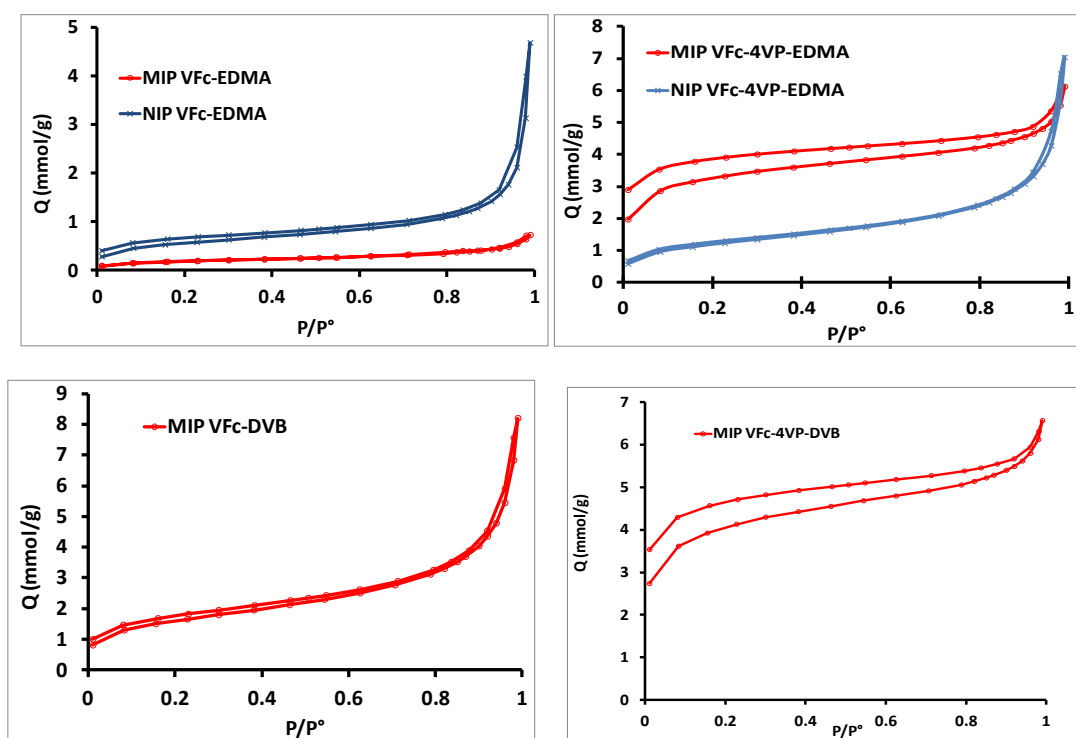


Figure 3. Nitrogen adsorption/desorption isotherms for e-MIP/e-NIP for VFc-EDMA and VFc-4VP-EDMA, e-MIP VFc-DVB and VFc-4VP-DVB.

### 3.2 Binding properties

The binding properties of BaP were investigated in an aqueous rich medium (water/acetonitrile (99:1, v/v)) in batch mode. Retention of BaP increased sharply in the early hours to lead at a maximum and constant after 6 hours contact time until the end of the experiment to 24 hours. Isotherms curves of e-MIP/e-NIP copolymers for VFc-EDMA, VFc-

DVB and VFc-4VP-DVB can be treated to type L isothermal (Figure 4) [18,19].

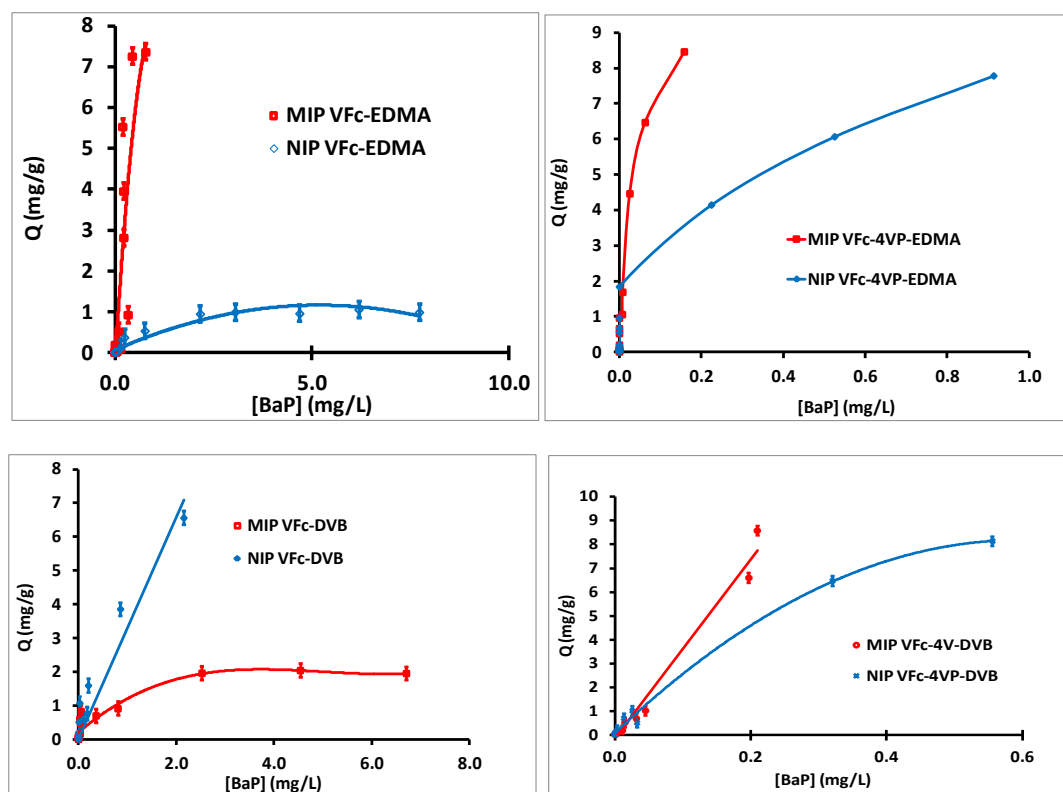


Figure 4. Adsorption isotherms of e-MIPs/e-NIPs for VFc-EDMA, VFc-4VP-EDMA, VFc-DVB and VFc-4VP-DVB depending on the equilibrium concentration ( $C_e$ ) of BaP in water/acetonitrile 99/1 (v/v).

In this case, the ratio between the adsorption capacity and the concentration of the remaining target (concentration at the equilibrium,  $C_e$ ) decreases when the target concentration increases. VFc-4VP-EDMA e-MIP and e-NIP presented H type isotherms which are a special case of L type isotherms having a very high initial slope. The low solubility of BaP in water/acetonitrile 99/1 water did not allow to reach the plateau. In these conditions, adsorption capacities were determined as the maximum values reached on isotherms and (Table 2) reveal similar values for e-MIPs and e-NIPs from VFc-4VP-DVB and VFc-4VP-EDMA with an imprinting factor close to 1.

Table 2. Values of adsorption capacity ( $Q$ ) and imprinting factors (IF) of the different polymers.

e-MIP Polymer	$Q$ (mg/g)	e-NIP Polymer	$Q$ (mg/g)	IF
---------------	------------	---------------	------------	----

VFc-EDMA	7.3±0.2	VFc-EDMA	1.0±0.2	7.3
VFc-4VP-EDMA	8.5±0.2	VFc-4VP-EDMA	7.8±0.2	1.1
VFc-DVB	2.0±0.2	VFc-DVB	6.3±0.2	0.3
VFc-4VP-DVB	8.6±0.2	VFc-4VP-DVB	8.1±0.2	1.1

Values for e-MIP/e-NIP from VFc-DVB copolymers indicated that e-NIP adsorb much more BaP compared to corresponding e-MIP revealing in this case that the synthesis of e-MIP was not needed to increase the efficiency of adsorption of BaP. The best imprinting effect was obtained for e-MIP-VFc-EDMA with a value of 7.3. Comparison of result between the use of EDMA and DVB were consistent with studies of W. L. Ho et al [20]. These authors compared MIP based on 4VP-EDMA and 4VP-DVB for BaP and found a better adsorption capacity for MIP-4VP-DVB (20.78 µg/g, IF 1.34) but a better imprinting effect for MIP-4VP-EDMA (4.77 µg/g, IF 2.30). As expected the replacement of part of VFc by 4VP increased the adsorption capacity. Unfortunately, there was no improving of imprinting factor.

Cross-selectivity of e-MIPs towards other PAHs was studied by putting the imprinted and non-imprinted polymers in contact with solutions of BaP with one interfering PAH chosen between fluorene, phenanthrene, anthracene, fluoranthene, pyrene, chrysene or dibenzo(a,h)anthracene). Based on the BaP rebinding results, the study of cross-selectivity was made only on e-MIP/e-NIP VFc-EDMA copolymers. The selectivity coefficients (k) ranged from 0.03 to 2.60 for e-MIP and from 0.02 to 4.30 for e-NIP (Table 3).

Table 3. Selectivity coefficients (k) and relative selectivity coefficients (k') for each PAH with VFc-EDMA copolymer.

	k	k	<i>k'</i>
	MIP	NIP	
Fluorene	0.11	0.15	<b>0.73</b>
Phenanthrene	1.33	0.02	<b>66.5</b>
Anthracene	0.03	0.63	<b>0.05</b>
Fluoranthene	1.39	0.76	<b>1.83</b>
Pyrene	0.90	0.47	<b>1.91</b>
Chrysene	0.38	0.62	<b>0.61</b>
Dibenzo(a,h)anthracène	2.60	4.30	<b>0.60</b>

For e-MIP-VFc-EDMA the low values of selectivity coefficients (k) indicated a poor recognition effect of BaP compared to other PAHs. The low range of k values is in agreement



with the very similar aromatic structures of BaP and other PAHs, and the lack of specific chemical functionalities for such compounds. For phenanthrene, fluoranthrene and pyrene the relative selectivity coefficients  $k'$ , defined as the ratio of  $k(\text{e-MIP})/k(\text{e-NIP})$ , were in favor of the imprinted polymers with an important effect in the case of phenanthrene. This still showed the imprinting effect and spoke in favor of specific interactions between BaP and the vinylferrocene monomer.

### 3.3 Benzo[a]pyrene electrochemical detection

Carbon paste was chosen to immobilize e-MIP and e-NIP for electrochemical characterizations. By this way, it was very simple and easy to fabricate modified electrode and it presented other advantages such as stability in various solvents, relatively low-background characteristics and a low residual current [21]. Only e-MIP/e-NIP based on VFc-EDMA and VFc-4VP-EDMA were tested. For each polymer, a mixture of e-MIP or e-NIP with carbon paste (CPE) (1:4, w/w) was prepared and put in a 5 mm diameter empty working electrode. After drying under vacuum, the carbon paste custom working electrode was polished and used in a classical three electrodes cell equipped with platinum as counter electrode and Saturated Calomel Electrode (SCE) as reference.

As expected anodic SWV scans for e-MIP-CPE and e-NIP-CPE based on VFc-EDMA showed one electron oxidation peak of ferrocene at 0.42 and 0.43V vs SCE, respectively (Figure 5). For e-MIP-CPE/e-NIP-CPE based on VFc-4VP-EDMA a positive shift of around 100 mV of oxidation peak was observed compared to VFc-EDMA electrode and could be attributed to an electron withdrawing effect of 4VP on ferrocenyl group by pi-pi interactions. Current intensity of anodic peaks for VFc-4VP-EDMA-CPE polymers was less intense in good agreement with the less quantity of ferrocenyl groups in these copolymers. In all cases, addition of BaP in the electrolytic solution leads to a decrease of current intensity with a more important impact for copolymers based on VFc-EDMA.



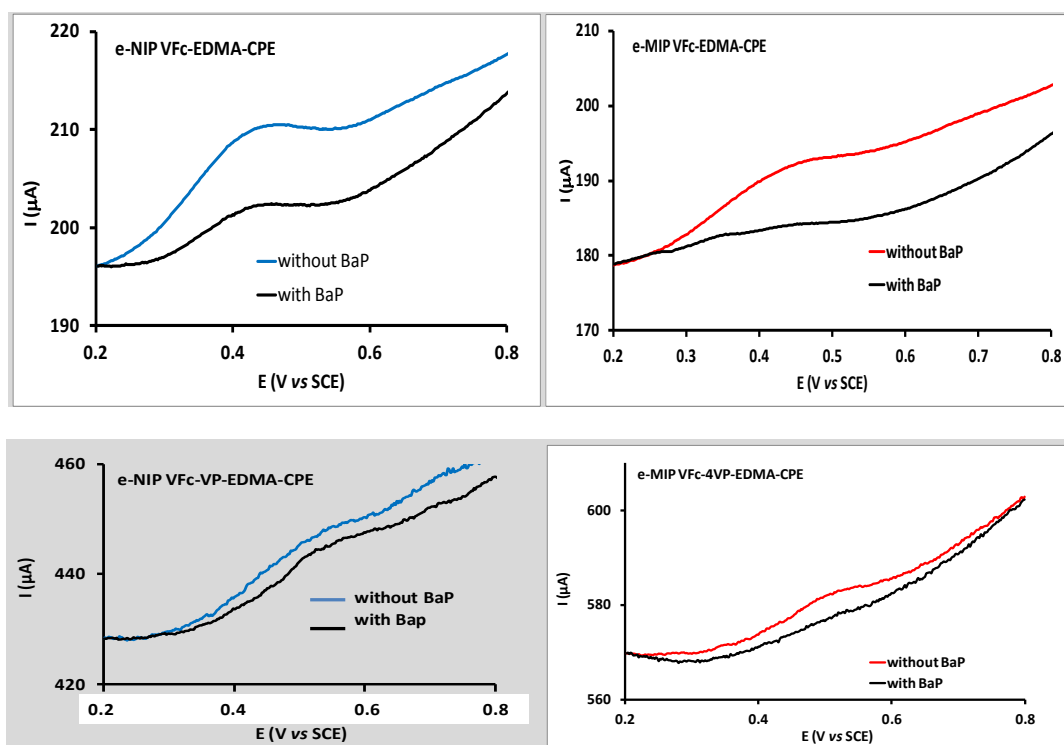


Figure 5. SWV of e-MIP-CPE/e-NIP-CPE with and without BaP (1 mg/L) after 4h. (a) e-MIP VFc-EDMA-CPE, (b) e-NIP VFc-EDMA-CPE, (c) e-MIP VFc-4VP-EDMA-CPE and (d) e-NIP VFc-4VP-EDMA-CPE

Working electrodes with e-MIP or e-NIP were immersed in 5 mL of electrolyte solution in the presence of increasing concentration of BaP from 0 to 4.5 mg/L (18  $\mu\text{mol/L}$ ) for 4 hours to reach adsorption equilibrium and the response of the electrode was registered (Figure 6). The intensity of the oxidation current was compared with that obtained in the absence of BaP. The increase of the normalized value of the oxidation peak current was proportional to the concentration of BaP added in the solution ( $R^2 > 0.989$ ). The slope of this linear function was 1.12  $\mu\text{mol/L}/\%$  for the e-MIP VFc-EDMA-CPE and 0.11  $\mu\text{mol/L}/\%$  for the e-MIP VFc-4VP-EDMA-CPE indicating that the first one is more sensitive to the presence of BaP. For both couples, e-MIP-CPEs had higher sensitivity in the presence of BaP than the corresponding e-NIPs-CPE. The ratio slope e-MIP-CPE/slope e-NIP-CPE was 1.4 for VFc-EDMA-CPE and 1.9 for VFc-4VP-EDMA-CPE. The limits of detection calculated for e-MIP VFc-EDMA-CPE and e-NIP VFc-EDMA-CPE were found to be 0.28 and 0.29 mM of BaP, respectively. The same calculation for e-MIP VFc-4VP-EDMA-CPE and e-NIP VFc-4VP-EDMA-CPE lead to 0.74 and 0.78  $\text{mmol/L}$  of BaP, respectively. Limits of quantification obtained for e-MIP VFc-EDMA-CPE, e-NIP VFc-EDMA-CPE, e-MIP VFc-4VP-EDMA-CPE and e-NIP VFc-4VP-EDMA-CPE were 0.93, 0.95, 2.47 and 2.59  $\mu\text{mol/L}$ , respectively.

Adsorption studies have revealed a maximum adsorption capacity similar for both polymers e-MIP VFc-EDMA and e-MIP VFc-4VP-EDMA (7.3 and 8.5 mg/g, respectively) but when incorporated inside electrodes, electrochemical analyses proved that e-MIP VFc-EDMA-CPE and corresponding e-NIP was more sensitive to the presence of BaP compared to polymers based on VFc-4VP-EDMA. These results were in good agreement with the amount of integrated vinylferrocene in these polymers. Indeed, the VFc was present at 18 to 19 mol% in the VFc-EDMA, unlike VFc-4VP-EDMA in which it was present only to 8% molar. The amount of VFc had a direct impact on the oxidation current value of the polymer, the higher its value, the greater the oxidation current was important.

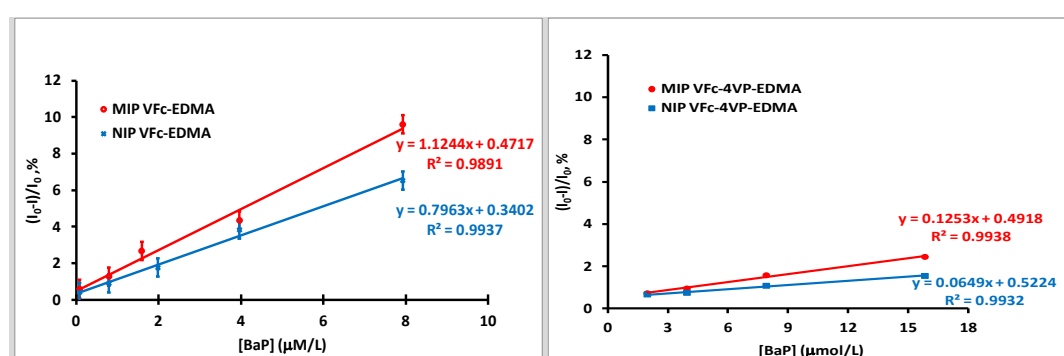


Figure 6. Normalized current curves where  $I$  was the current obtained in presence of BaP and  $I_0$  the current before BaP addition (formulated in percentage in function of added BaP). (a) VFc-EDMA-CPE et (b) VFc-4VP-EDMA.

#### 4. Conclusion

In conclusion, original electrochemical molecularly imprinted polymers (e-MIPs) with redox binding cavities to detect benzo(a)pyrene (BaP) were synthesized integrating the vinylferrocene (VFc) redox tracer with or without 4-vinylpyridine (4-VP) as comonomer and ethylene glycol dimethacrylate (EDMA) or divinylbenzene (DVB) as crosslinker, in the same time with a hybridation oven. As expected, using an aromatic comonomer or crosslinker increase adsorption capacities by a factor 6-8 compared to the reference polymer VFc-EDMA but, unfortunately, not the imprinting factor. Electrochemical measurements for e-MIP/e-NIP for polymers based on VFc-EDMA and VFc-4VP-EDMA revealed that the first one has the best LOD with a value of  $0.93 \mu\text{mol/L}$ . Knowing that these carbon paste electrodes are not optimized, the strategy reported in the present article is very promising for future applications in PAHs sensors fields and can be, moreover, applied to a wild range of analytes presenting or lacking redox properties.

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